Advances in Process Simulation of Agglomeration in Bayer Precipitation

Ab Rijkeboer¹, John McFeaters² and Daniel Manché³

1. Director, Rinalco BV, Wassenaar, The Netherlands

- 2. Director, US Office, KWA Kenwalt Australia Pty Ltd, Baton Rouge, LA, USA
- 3. Simulation Engineer, Europe Office, KWA Kenwalt Australia Pty Ltd, Lisbon, Portugal Corresponding author: ab@rinalco.com

Abstract



Developing a Bayer precipitation circuit model with accurate predictive capability requires comprehensive physical models for the fundamental processes occurring and, importantly, a structured approach to tuning the model to available plant data. A new agglomeration size kernel has been constructed based on agglomeration kernels developed by René David et al. for different types of eddies within turbulence. The new kernel takes account of the fact that two types of eddies are of practical relevance to agglomeration in precipitation circuits, namely small laminar eddies whose behaviour is dominated by viscous forces and larger eddies whose behaviour is dominated by inertial forces. As well as simulating the chemical and physical processes underlying agglomeration, the kernel offers flexibility in its application due to the inclusion of various adjustable parameters. This facilitates model calibration for accurate prediction of particle size distribution (PSD) in the discharge flows from tanks in the precipitation circuit where complex particle interactions take place under a variety of hydrodynamic conditions. The kernel has been implemented into the Precipitator 3 unit model of the SysCAD process simulation software. A systematic procedure for calibration has been developed for obtaining the closest approach to the target PSD whereby various combinations of kernel parameters are tested with simultaneous tuning of constants for growth, nucleation, and agglomeration rates on their respective targets. Aspects of supersaturation, growth rate, nucleation rate and agglomeration rate are reviewed in the context of calibrating predictive established relationships.

Keywords: Process simulation, Agglomeration, Nucleation, Growth, Calibration.

1. Introduction

Product quality and production rate are ongoing concerns in alumina refineries. Uncertainties often arise about the best way to resolve product quality issues with the least impact on production or how to optimise plans for production improvements. In these cases it is useful to have a tool which can predict changes in particle size distribution (PSD) around the white-side circuit as a consequence of changes in operating parameters, circuit configuration or equipment conditions.

Recently a new agglomeration size kernel has been added to the SysCAD process simulation software. The kernel offers a large degree of flexibility in calibrating a precipitation circuit model to an existing plant circuit such that each tank in the model has the right set of parameters producing the correct gibbsite mass with the correct PSD. This paper discusses this new kernel and the method of calibration. In this context, attention is also paid to fundamental aspects of growth rate, nucleation rate and agglomeration rate.

Agglomeration of Al(OH)₃ (gibbsite or hydrate) crystals in the Bayer process has been extensively studied in Australian research institutions. In the early 2000 s a collaborative research project was executed within CSIRO Minerals under the direction of AMIRA. Dean Ilievski and Iztok Livk were the leading scientists on this project and published important results, conclusions and correlations, including an agglomeration kernel that is available in SysCAD. Their experimental

work was conducted in two types of reactors: an in-house developed reactor for simulating laminar flow conditions and another for simulating turbulent flow conditions [1]. Regarding the latter it was recognised that the agitator power input, thus shear rate, per reactor volume unit was an order of magnitude greater than in real-world precipitators. To overcome this issue, the conclusions from both the laminar and turbulent reactors were considered in constructing an agglomeration kernel.

In this paper a more pragmatic approach is presented whereby the actual precipitators in a plant serve as kernel parameter development reactors. In this approach, scientific findings regarding agglomeration under laminar and turbulent flow conditions are used, but a degree of flexibility has been incorporated into the kernel to account for the actual plant environment.

2. New Agglomeration Kernel

2.1 Microscales of Turbulence

The new kernel is based on expressions from publications by René David et al. [2, 3, 4] about their research into the agglomeration of adipic acid crystals. These expressions relate to the hydrodynamic and orthokinetic conditions within the turbulence in a tank, where different types of eddies can be distinguished. Turbulent energy is added at the largest length scales. The vortices or eddies created are distorted and broken into continually smaller eddies. Of interest here are eddies at the smaller scales, where agitator energy input is ultimately dissipated as heat via viscous action. The Taylor microscale, λ_g , is the scale at which viscous effects start to become important, while turbulent inertia through flow fluctuations in magnitude and direction is still present. Below the Kolmogorov microscale, λ_k [5], viscous effects dominate and motion is laminar.

$$\lambda_k = \left(\frac{\nu^3}{\varepsilon}\right)^{1/4} \tag{1}$$

$$\lambda_g = \theta \cdot \left(\frac{60\nu}{\varepsilon}\right)^{1/2} \tag{2}$$

where:

λ_k	Kolmogorov microscale of turbulence, m
λ_g	Taylor scale of turbulence, m
ν	Kinematic viscosity, m ² s ⁻¹
ε	Energy dissipation rate per unit mass, J s ⁻¹ kg ⁻¹
θ	Fluctuating average component of the local velocity, m s ⁻¹

The Taylor and Kolmogorov microscales are typically within the hydrate particle size range of Bayer circuit precipitators. These scales can vary from tank to tank and from location to location within a tank. It is worth noting that turbulence is a characteristic of the flow situation, not a property of the fluid, and that these scales are not exact dimensions but representative dimensions.

2.2 The Process of Agglomeration

The Kolmogorov microscale is an important transition point for particle behaviour. The motion of particles smaller than the Kolmogorov microscale is dominated by viscous forces. These particles have little motion relative to the fluid. They tend to move parallel to each other in the laminar flow elements. Collisions between these small particles only occur because particles catch up with each other as a result of a viscosity related, shear induced velocity gradient. The chance of collision as a function of particle size is described by Smoluchovski [6]. Particles larger than the Kolmogorov microscale tend to follow more diverse individual flow patterns and exhibit varied velocity fluctuations and trajectory changes. They have a much greater chance of being involved in collisions, e.g. with each other, or with the smaller particles crossing their path.

7. References

- 1. I.Livk, D.Ilievski, A macroscopic agglomeration kernel model for gibbsite precipitation in turbulent and laminar flows, *Chemical Engineering Science*, 62 (2007), 3787-3797.
- 2. René David et al., Modelling of multiple-mechanism agglomeration in a crystallization process, *Powder Technology, Elsevier*, (2003), 338-344.
- 3. René David et al., Developments in the understanding and modeling of the agglomeration of suspended crystals in crystallization from solutions, *Kona*, No.21 (2003), 40-53.
- 4. René David et al, Crystallization and precipitation engineering III. A discrete formulation of the agglomeration rate of crystals in a crystallization process, *Chemical Engineering Science*, Vol.46 (1991), 205-213.
- 5. Andrey Kolmogorov, Dissipation of energy in locally isotropic turbulence, *Doklady Academy* of Science, USSR (1941), 16-18.
- 6. M.V.Smoluchowski, Versuch einer mathematischen Theorie der Koagulationskinetik Kolloider Lösungen, Zeitschrift fur Physikalische Chemie, (1917), 129-168.
- 7. Mei-yin Lee, Gordon M. Parkinson, Growth rates of gibbsite single crystals determined using in situ optical microscopy, *Journal of Crystal Growth*, (1999), 270-274.
- 8. C.Misra, E.T.White, Kinetics of Crystallisation of aluminium trihydroxide from seeded caustic aluminate solutions, *Chem. Eng. Prog. Symposium*, Series 67 (1971), 53-65.
- 9. D.Ilievski, E.T.White, Agglomeration during precipitation: Agglomeration mechanism identification for Al(OH)₃ crystals in stirred caustic aluminate solutions, *Chemical Engineering Science*, (1994), 3227-3239.
- 10. M.J.Hounslow, R.L.Ryall, V.R.Marshall, A discretized population balance for nucleation, growth, and aggregation, *AICHE Journal*, Vol.34, No.11 (November 1988), 1821-1832.
- 11. R.R.M.Johnston, P.J.Cresswell, Modelling alumina precipitation: dynamic solution of the population balance equation, *Proc. Fourth Alumina Quality Workshop*, (1996), 282-290.
- 12. Rod Stephenson, A quasi-implicit algorithm for solution of steady state and dynamic particle balance models, *Unpublished paper from poster session at the Ninth Alumina Quality Workshop*, (2012), available on request.
- 13. O.Tschamper, Improvements by the new Alusuisse process for producing coarse aluminum hydrate in the Bayer process, *Proc. AIME Light Metals Conference*, (1981), 103-115.
- 14. V.G.Levich, The theory of coagulation of colloids in turbulent liquid stream, *Dokl. Akad. Nauk SSSR*, Vol.99 (1954), 809-812.
- 15. D.N.Seneviratne et al, An investigation of Bayer desilication product agglomeration mechanism by kernel function population balance modelling, *Proc. Eleventh Alumina Quality Workshop*, (2018).
- 16. T.S.Li, I.Livk, D.Ilievski, Supersaturation and temperature dependency of gibbsite growth in laminar and turbulent flows, *Journal of Crystal Growth*, 258 (2003), 409-419.
- 17. H.R.Watling et al, Ionic structure in caustic aluminate solutions and the precipitation of gibbsite, *J. Chem. Soc. Dalton Trans.*, (1998), 3911-3917.
- 18. C.F.Vernon, G.M.Parkinson, D.Lau, Towards a fundamental rate equation for gibbsite growth in Bayer liquors, *Proc. Fifth Alumina Quality Workshop*, (1999), 129-139.
- 19. C.Sweegers et al., Growth rate analysis of gibbsite single crystals growing from aqueous sodium aluminate solutions, *Crystal Growth & Design*, Vol. 4, No. 1 (2004), 185-198.
- 20. R.M.Cornell et al., Precipitation of gibbsite: Development of a new rate equation, *Proc. Fifth Alumina Quality Workshop*, (1999), 153-161.
- 21. Dean Ilievski, Can gibbsite precipitation be modelled from first principles ?, *Proc. Fourth Alumina Quality Workshop*, (1996), 282-290.
- 22. S.Veesler, R.Boistelle, About supersaturation and growth rates of hydrargillite Al(OH)₃, *Journal of Crystal Growth*, 130 (1993), 411-415.
- 23. S.Veesler, R.Boistelle, Growth kinetics of hydrargillite Al(OH)₃ from caustic soda solutions, *Journal of Crystal Growth*, 142 (1994), 177-183.

- 24. C.F.Vernon et al, Mechanistic investigations of gibbsite growth, *Proc. Sixth Alumina Quality Workshop*, (2002), 33-39.
- 25. H.R.Watling et al., IR and ²⁷Al-MAS-NMR spectroscopic studies of sodium(hydroxy)aluminates, *Applied Spectroscopy*, 53-4 (1999), 415-422.
- 26. Joanne Sook Ching Loh, The role of cations in gibbsite crystallization, *PhD Thesis, Curtin University of Technology*, (January 2001).
- 27. J.Addai-Mensah, C.A.Prestige, J.Ralston, Interparticle forces, interfacial structure development and agglomeration of gibbsite particles in synthetic Bayer liquors, *Minerals Engineering*, (1999), 655-669.
- 28. John Garside, Industrial crystallization from solution, *Chemical Engineering Science*, Vol. 40, No. 1 (1985), 3-26.
- 29. S.Friej et al, Investigation of the growth of gibbsite crystals by atomic force microscopy and optical microscopy, *Proc. Fifth Alumina Quality Workshop*, (1999), 153-161.
- 30. N.Brown, Secondary nucleation of aluminium trihydroxide in seeded caustic aluminate solutions, Journal of Crystal Growth, (1972), 163-169.
- 31. C.Sweegers et al, Surface Topography of Gibbsite Crystals grown from Aqueous Sodium Aluminate Solutions, Applied Surface Science, Vol. 187 No. 3-4 (2002), 218-234.
- 32. C.Misra, The precipitation of Bayer aluminium trihydroxide, PhD Thesis, University of Queensland, (September 1970).
- 33. M.M.Reyhani et al, Gibbsite nucleation at sodium oxalate surfaces, Proc. Fifth Alumina Quality Workshop, (1999), 181-191.
- 34. D.Ilievski, I.Livk, An agglomeration efficiency model for gibbsite precipitation in a turbulently stirred vessel, Chemical Engineering Science, (2006), 2010-2022.
- 35. D.Ilievski, Modelling Al(OH)3 agglomeration during batch and continuous precipitation in supersaturated caustic aluminate solutions, PhD Thesis, University of Queensland, (1991).
- 36. A.Halfon, S.Kaliaguine, Alumina trihydrate crystallisation. Part 1, Secondary nucleation and growth rate kinetics, Canadian Journal of Chemical Engineering, (1994), 160.
- 37. D.Green, R.Perry, Modelling and simulation of granulation processes, Perry's Chemical Engineers' Handbook, 8th Edition (2008).
- 38. S.C.Grocott, S.P.Rosenberg, Sodium in alumina. Possible mechanisms for soda incorporation, Proc. First Alumina Quality Workshop, (1988), 271-287.
- 39. J.V.Sang, Factors affecting residual soda in precipitation products, Proc. AIME Light Metals Conference, (1988), 147-156.
- 40. J.Ohkawa, T.Tsuneizumi, T.Hirao, Technology of controlling soda pick-up in alumina trihydrate precipitation, Proc. AIME Light Metals Conference, (1985), 345-366.
- 41. T.S.Li et al, The influence of solid phase oxalate (SPO) on gibbsite secondary nucleation in synthetic caustic-aluminate solution, Proc. Tenth Alumina Quality Workshop, (2015), 165-173.